



# Thermodynamic analysis of glycerol conversion to synthesis gas

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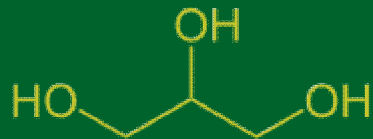
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# Why glycerol? Why synthesis gas?

## Glycerol



- unique structure and properties
- bioavailability
- renewability
  
- biodiesel production**  
(main side-product: ~ 10 wt. %)
- numerous possibilities for its conversion to:
  - acrolein and acetol
  - diols (1,2- and 1,3-propanediol, ethylene glycol)
  - glycerol carbonate
  - **synthesis gas**

## Synthesis gas



- ❖ many applications in chemical synthesis
  - methanol production
  - hydrogen production
  - oxosynthesis
  - Fischer-Tropsch process



# Thermodynamic analysis of glycerol conversion to synthesis gas

Main reagents of glycerol conversion (with or without CO<sub>2</sub>):

**Glycerol, carbon monoxide, carbon dioxide, hydrogen, water, methane**

**Table 1. Selected stoichiometric models.**

1 <sup>st</sup> model	$\text{C}_3\text{H}_8\text{O}_3 + 3\text{H}_2\text{O} = 3\text{CO}_2 + 7\text{H}_2 \text{ (1a)}$ $\text{CO}_2 + \text{H}_2 = \text{CO} + \text{H}_2\text{O} \text{ (1b)}$ $\text{CO}_2 + 4\text{H}_2 = \text{CH}_4 + 2\text{H}_2\text{O} \text{ (1c)}$
2 <sup>nd</sup> model	$\text{C}_3\text{H}_8\text{O}_3 + \text{CO}_2 = 4\text{CO} + 3\text{H}_2 + \text{H}_2\text{O} \text{ (2a)}$ $\text{CO}_2 + \text{H}_2 = \text{CO} + \text{H}_2\text{O} \text{ (2b)}$ $\text{CO} + 3\text{H}_2 = \text{CH}_4 + \text{H}_2\text{O} \text{ (2c)}$
3 <sup>rd</sup> model	$\text{C}_3\text{H}_8\text{O}_3 = 3\text{CO} + 4\text{H}_2 \text{ (3a)}$ $\text{CO} + 3\text{H}_2 = \text{CH}_4 + \text{H}_2\text{O} \text{ (3b)}$ $\text{CO} + \text{H}_2\text{O} = \text{CO}_2 + \text{H}_2 \text{ (3c)}$

**Glycerol conversion with steam**  
*(steam reforming)*

**Glycerol conversion with carbon dioxide**  
*(dry reforming)*

**Glycerol conversion without any co-substrate**

## Conclusions



**It was found that the contribution of synthesis gas in the equilibrium mixture increases with the increase temperature.**



**In the steam reforming of glycerol synthesis gas rich with hydrogen is produced, while in the dry reforming the carbon monoxide predominates in the equilibrium mixture.**

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